

**Notes:**

In chemistry an intensive property of a system is a physical property of the system that does not depend on the system size or the amount of material in the system. By contrast, an extensive property of a system does depend on the system size or the amount of material in the system.

Examples of intensive properties include:

- \* temperature
- \* viscosity
- \* density
- \* electrical resistivity
- \* melting point
- \* boiling point
- \* pressure
- \* spectral absorption maxima (in solution)
- \* flammability

Examples of extensive properties include:

- \* mass
- \* volume
- \* entropy
- \* energy
- \* electrical resistance
- \* texture
- \* heat

Also intensive property does not depend on the physical quantity. For example consider one liter of water at 20 degrees. Now even if you make the water half a liter the temperature is going to remain the same. so temperature is an intensive property. but when u make it half a liter the volume decreases thus making volume an extensive property as it depends on the amount.